

**CHEMISTRY NMDCAT****SWIFT TEST**

- Q.1 Which of the following enthalpies of formation cannot be measured directly
 (a) $\Delta H^\circ_{\text{ion}}$ for ionic compound (b) ΔH°_f for $\text{CO}_{(g)}$
 (c) ΔH°_f for B_2O_3 (d) All of these
- Q.2 Which is a non-spontaneous reaction
 (a) $\text{RbOH} + \text{HNO}_3 \rightarrow \text{RbNO}_3 + \text{H}_2\text{O}$ (b) $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
 (c) $\text{N}_2 + \text{O}_2 \rightarrow 2\text{NO}$ (d) $\text{K} + \text{H}_2\text{O} \rightarrow \text{KOH} + \text{H}_2$
- Q.3 Which one of the following has maximum bond angle
 (a) CCl_4 (b) NF_3
 (c) NH_3 (d) CO_2
- Q.4 The correct order of electron affinity
 (a) $\text{F} > \text{Cl} > \text{I}$ (b) $\text{Cl} > \text{F} > \text{I}$
 (c) $\text{Cl} > \text{I} > \text{F}$ (d) $\text{I} > \text{Cl} > \text{F}$
- Q.5 A buffer solution has equal volume 0.5M NH_4OH and 0.5M NH_4Cl . The pK_b of base is 4.74. pH of the solution is
 (a) 10.26 (b) 9.26
 (c) 4.74 (d) 7.26
- Q.6 The solubility product of PbSO_4 is $4 \times 10^{-4} \text{ mol}^2 \cdot \text{dm}^{-4}$. The maximum concentration of Pb^{2+} ions is
 (a) $1.41 \times 10^{-3} \text{ mol} \cdot \text{dm}^{-3}$ (b) $2 \times 10^{-3} \text{ mol} \cdot \text{dm}^{-3}$
 (c) $2 \times 10^{-3} \text{ mole} \cdot \text{dm}^{-3}$ (d) $4 \times 10^{-3} \text{ mol} \cdot \text{dm}^{-3}$
- Q.7 Half-life of a 1st order reaction $\text{A} \rightarrow \text{B}$ is $\left(\frac{0.693}{k}\right)$ seconds its rate constant is
 (a) 0.231 s^{-1} (b) 2 s^{-1}
 (c) 4 s^{-1} (d) 20 s^{-1}
- Q.8 The rate of reaction can be increased in general by all of the following factors except
 (a) By increasing the temperature (b) Using a suitable catalyst
 (c) By increasing activation energy (d) By increasing concentration of reactants
- Q.9 Following first order reaction is 50 percent completed in 24 minutes at 300K
 $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$ [N₂O₅] given = 10g
 How many grams of N_2O_5 will be left behind after 72 minutes?
 (a) 1.77g (b) 1.25g
 (c) 2.5g (d) 0.630g
- Q.10 Which of the following is the least reactive alkali metal
 (a) Lithium (c) Sodium
 (b) Potassium (d) Cesium
- Q.11 When Ba is heated in air or oxygen at 500–600°C temperature, it is converted into
 (a) Normal oxide (b) Peroxide
 (c) Superoxide (d) Suboxide
- Q.12 Which enzyme bring about exchange in functional group between two compounds
 (a) Phospho-transferase (b) Phospho-glyceromutase
 (c) L-asparaginase (d) LDH-1
- Q.13 Many enzymes contain a protein part and non-protein part. This protein part is _____
 (a) Apoenzyme (b) Holoenzyme
 (c) Co-factor (d) Co-enzyme
- Q.14 Phosphoprotein comes under the type of proteins
 A) Simple protein C) Derived protein
 B) Conjugated D) Both A & B
- Q.15 Caloric value of fuel can be determined by
 (a) Copper calorimeter (b) Glass calorimeter
 (c) Bomb calorimeter (d) Flame calorimeter
- Q.16 Stronger is the oxidizing agent, greater is the
 (a) Oxidation potential (b) Reduction potential
 (c) Standard reduction potential (d) Redox potential



- Q.17 Generally, when a metal reacts with halogens the decreasing order of ionic character of compounds is
 (a) $\text{NaF} > \text{NaCl} > \text{NaBr} > \text{NaI}$ (b) $\text{NaF} > \text{NaBr} > \text{NaCl} > \text{NaI}$
 (c) $\text{NaI} > \text{NaBr} > \text{NaCl} > \text{NaF}$ (d) $\text{NaCl} > \text{NaF} > \text{NaBr} > \text{NaI}$
- Q.18 X, Y and Z are elements in the same short period. The oxide of "X" is amphoteric, oxide of "Y" is basic and oxide of "Z" is acidic. What is the order of increasing atomic number these elements
 (a) X-Y-Z (b) Y-X-Z
 (c) X-Z-Y (d) Y-Z-X
- Q.19 The rate constant depends upon
 (a) Half life period (b) Activation energy
 (c) Temperature (d) Both B and C
- Q.20 When temperature of an exothermic reaction is increased then
 (a) Rate of reaction is increased (b) Rate of reaction is decreased
 (c) Energy of activation is decreased (d) Yield is increased
- Q.21 Which one is correct relation
 (a) $\text{pOH} = \text{pKa} + \log \frac{[\text{Salt}]}{[\text{Base}]}$ (b) $\text{pOH} = \text{pKb} + \log \frac{[\text{Salt}]}{[\text{Acid}]}$
 (c) $\text{pOH} = \text{pKb} + \log \frac{[\text{Base}]}{[\text{Salt}]}$ (d) $\text{pOH} = \text{pKb} + \log \frac{[\text{Salt}]}{[\text{Base}]}$
- Q.22 The first four ionization energies values of an element are 284, 412, 656 and 3210 kJmol^{-1} . The number of valence electrons in the element are
 (a) One (b) Three
 (c) Two (d) Four
- Q.23 The correct decreasing order of size among the following species
 (a) $\text{Ar} > \text{Cl}^- > \text{S}^{2-}$ (b) $\text{Cl}^- > \text{Ar} > \text{S}^{2-}$
 (c) $\text{S}^{2-} > \text{Cl}^- > \text{Ar}$ (d) $\text{S}^{2-} > \text{Ar} > \text{Cl}^-$
- Q.24 Element that shows peculiar behavior among alkaline earth metals is
 (a) Be (b) Mg
 (c) Ba (d) Ca
- Q.25 Which of the following is correct data about $\text{K}_4[\text{Fe}(\text{CN})_6]$?
- | | O.N of Central Metal | Coordination No. |
|---|----------------------|------------------|
| A | +2 | 4 |
| B | +1 | 2 |
| C | +2 | 6 |
| D | -2 | 6 |
- Q.26 The number of lone pair of electrons provided by the ligands to the central metal atom or ion is called
 (a) Oxidation number (b) Coordination number
 (c) Coordination sphere (d) Charge on sphere
- Q.27 Which of the following is more basic oxide of group IIA elements
 (a) BeO (b) MgO
 (c) SrO (d) CaO
- Q.28 _____ is used as catalyst in Haber's process for NH_3 gas manufacturing
 (a) Iron (b) Copper
 (c) Carbon (d) Silver
- Q.29 Acetone and propen-2-ol are
 (a) Positional isomers (b) Keto-enol tautomers
 (c) Geometrical isomers (d) Chain isomers
- Q.30 The IUPAC name of $(\text{C}_2\text{H}_5)_2\text{C}(\text{CH}_3)_2$ is
 (a) 1,1-Diethyl-1,1-dimethylmethane (b) 3,3-Dimethylpentane
 (c) 2,2-Dimethylpropane (d) 2,2-Diethylpropane
- Q.31 Which of the following act as nucleophile in aldol condensation reaction
 (a) Hydroxide ion (b) Carbanion ion
 (c) Hydride ion (d) Carbocation ion



- Q.32 In the addition of HBr to butene, the first step involves the addition of
(a) Br⁻ (b) Br⁺
(c) Br⁻ (d) H⁺
- Q.33 The most reactive species towards electrophilic substitution reactions is
(a) Aniline (b) Chlorobenzene
(c) Benzene (d) Nitrobenzene
- Q.34 During chlorination of methane, if chlorine is used in excess, then major product is
(a) CH₃Cl (b) CHCl₃
(c) CH₂Cl₂ (d) CCl₄
- Q.35 Addition of which reagent to unsymmetrical alkenes obey Markownikoff rule
(a) HCl (b) HOCl
(c) HCN (d) All of these
- Q.36 $\text{CH}_3\text{CH}_2\text{Br} + \text{KOH} \xrightarrow{\text{alc}} \text{C}_2\text{H}_4 + \text{KBr} + \text{H}_2\text{O}$
The above reaction follow the mechanism
(a) S_N1 (b) S_N2
(c) E2 (d) E1
- Q.37 Which of the following is an element as well as covalent solid?
(a) Silica (b) Iodine
(c) Mercury (d) Diamond
- Q.38 Acetone and chloroform are soluble into each other due to
(a) Dipole-dipole interaction (b) Debye forces
(c) Intermolecular hydrogen bonding (d) London dispersion forces
- Q.39 Total number of fundamental particles in one atom of $^{14}_7\text{N}^{3-}$
(a) 7 (b) 15
(c) 8 (d) 24
- Q.40 Which of the following molecule has net dipole moment
(a) CO₂ (b) BF₃
(c) PF₃ (d) SO₂
- Q.41 Which of the following pair has same type of hybridization on central atom
(a) H₂S, SO₂ (b) BF₃, NH₃
(c) BF₃, SO₂ (d) CH₃, BF₃
- Q.42 Work done by the system is 50kJ, heat absorbed by the system is 75kJ, the ΔE will be
(a) +125kJ (b) -125kJ
(c) +25kJ (d) -25kJ
- Q.43 One of the followings is always exothermic
(a) Enthalpy of formation (b) Enthalpy of atomization
(c) Enthalpy of neutralization (d) Enthalpy of solution
- Q.44 The conductivity through external wire in galvanic cell is from
(a) Negatively charged anode to positively charged cathode
(b) Positively charged anode to negatively charged cathode
(c) Negatively charged cathode to positively charged anode
(d) Positively charged cathode to negatively charged anode
- Q.45 Down the group, oxidation state of representative elements
(a) Increases (b) Increases and then decreases
(c) Decreases (d) Remains same
- Q.46 Which of the following has lowest pK_a value
(a) Phenol (b) Water
(c) Acetic acid (d) Ethyl alcohol
- Q.47 Which of the following statement about K_c is incorrect?
(a) K_c value is only changed by the change of temperature
(b) It may or may not has unit
(c) Its value is independent of the initial conc. of reactants
(d) K_c is independent of temperature
- Q.48 The rate equation for a gas phase reaction is rate = k [A][B]. If pressure on reaction mixture is tripled, then the rate of reaction increases by
(a) Two times (b) Nine times
(c) Six times (d) Eight times
- Q.49 The diamminesilver(I) complex in the mixture of aldehyde and Tollen's reagent act as an oxidizing agent and its formula is
(a) $[\text{Ag}(\text{NH}_3)_2]^+$ (b) $[\text{Ag}(\text{NH}_3)_3]^+$
(c) $[\text{Ag}(\text{NO}_2)]^+$ (d) $[\text{Ag}(\text{NH}_3)_2]^{+2}$



- Q.50 The reaction of C_2H_5OH with Na metal is example of _____ reaction
(a) Nucleophilic addition (b) Electrophilic addition
(c) Nucleophilic substitution (d) Electrophilic substitution
- Q.51 Acetone on oxidation gives
(a) Propanoic acid (b) Acetic acid + Formic acid
(c) Acetic acid only (d) Acetic acid + Formaldehyde
- Q.52 The IUPAC name of Diethyl ketone
(a) 2-Butanone (b) 2-Pentanone
(c) 3-Pentanone (d) 3-Hexanone
- Q.53 The number of carbon atoms in cyclic dimer of a carboxylic acid are
(a) 2 (b) 4
(c) 3 (d) 6
- Q.54 An ester with raspberry flavour can be produced by reacting isobutyl alcohol with
(a) Formic acid (b) Butyric acid
(c) Acetic acid (d) Propionic acid
- Q.55 Acetic acid is more acidic than phenol due to
(a) Resonance stability of phenoxide ion
(b) Resonance stability of acetate ion
(c) Delocalization of π -electrons in phenoxide ion
(d) Lower pK_a value of phenol
- Q.56 The C - C bond distance is longest in
(a) C_2H_2 (b) C_2H_4
(c) C_2H_6 (d) C_6H_6
- Q.57 The nucleophile is
(a) $AlCl_3$ (b) H_3O^+
(c) NH_4^+ (d) Cl^-
- Q.58 Which of the following reagent is used for conversion of carboxylic acid into alkane
(a) $LiAlH_4$ /Ether (b) $NaBH_4/H_2$
(c) $KMnO_4/H_2SO_4$ (d) H/P
- Q.59 Atomic radius decreases from left to right in the period due to
(a) Increasing shielding effect (b) Increasing number of shells
(c) Decreasing nuclear charge (d) Increasing nuclear charge
- Q.60 The highest electrical conductivity is shown by
(a) Mg (b) Si
(c) Al (d) P
- Q.61 The correct electronic configuration of Cr ($Z = 24$) is
(a) $[Ne]4s^13d^5$ (b) $[Ne]4s^23d^4$
(c) $[Ar]4s^13d^5$ (d) $[Ar]4s^23d^4$
- Q.62 The toughness of transition metals indicates strong metallic binding. The binding is stronger upto group _____.
(a) IIB (b) VIB
(c) IIIB (d) IB
- Q.63 Which of the following shows peculiar behaviour in many of its properties
(a) Li (b) Na
(c) K (d) Rb
- Q.64 Incorrect statement is
(a) Nucleophile may be negative or neutral specie
(b) Electrophile may be positive or neutral specie
(c) Free radical may be charged or neutral specie
(d) Free radical always neutral specie
- Q.65 In benzene substituted compounds, the 2,4-directing group is
(a) $-COOH$ (b) $-CHO$
(c) $-NHR$ (d) $-NR_2$
- Q.66 Which of the followings can be used to distinguish between alkane and alkene
(a) $Aq. Br_2$ (b) $Alc. KOH$
(c) Alkaline $KMnO_4$ (d) Both "a" and "c"
- Q.67 Toluene when treated with acidified $KMnO_4$ produces
(a) Phenol (b) Methyl Cyclohexane
(c) Benzoic Acid (d) Glyoxal



- Q.68 Which would undergo S_N2 reactions faster
 (a) Methyl iodide (b) Ethyl iodide
 (c) Ethyl chloride (d) Propyl bromide
- Q.69 t -butyl bromide $\xrightarrow{alc.KOH}$ Z, Here 'Z' is
 (a) Propene (b) 2-Butene
 (c) Butene (d) 2-Methylpropene
- Q.70 Alkyl halide gives elimination reaction with alcoholic KOH and they give _____ reaction with alcoholic KCN
 (a) Nucleophilic substitution (b) Elimination
 (c) Electrophilic substitution (d) Nucleophilic addition
- Q.71 The factor that does not effect the rate of S_N1 reaction
 (a) Concentration of $R-X$ (b) Nature of Alkyl group
 (c) Concentration of Nucleophile (d) Nature of solvent
- Q.72 The IUPAC name of $HOOC-(CH_2)_3-COOH$ is
 (a) 1,5-Pentanoic acid (b) Pentanedioic acid
 (c) Pentane-1,5-dioic acid (d) Pentane-1,5-dial
- Q.73 The alcohol which readily gives electrophilic substitution reaction
 (a) Primary alcohols (b) Tertiary alcohols
 (c) Secondary alcohols (d) Methyl alcohol
- Q.74 The reaction of alcohol in which its $C-O$ bond is not broken
 (a) Alkyl halide formation (b) Ester formation
 (c) Alkene formation (d) None of these
- Q.75 Formation of yellow precipitate with $I_2/NaOH$ in case of alcohols indicate
 (a) Methanol (b) Propanol
 (c) Ethanol (d) Butanol
- Q.76 Alcohols can be converted into aldehydes and ketones by
 (a) $K_2Cr_2O_7 + H_2SO_4$ (b) $NaBH_4$
 (c) $LiAlH_4$ (d) $NaOH + Ni - Al$
- Q.77 Which of the following is not a mild oxidizing agent
 (a) Lucas reagent (b) Fehling's solution
 (c) Tollen's reagent (d) Benedict's solution
- Q.78 Which element has highest %age of oxygen?
 (a) CH_3OH (b) H_2O
 (c) CO (d) CH_2O
- Q.79 Empirical formula of glucose is
 (a) CH_2O (b) C_2H_2O
 (c) CHO_2 (d) CHO
- Q.80 Total number of covalent bonds present in 8g Methane are
 (a) 6.02×10^{23} (b) 1.2×10^{24}
 (c) 1.2×10^{23} (d) 1.5×10^{23}
- Q.81 A 6 g sample of organic compound produces 2g CO_2 and 1g H_2O during combustion. What is the %age of carbon
 (a) 40% (b) 33%
 (c) 9% (d) 15%
- Q.82 Absorption of H_2O and CO_2 in $Mg(ClO_4)_2$ and 50% KOH are _____ and _____ process respectively
 (a) Physical, Physical (b) Physical, Chemical
 (c) Chemical, Physical (d) Chemical, Chemical
- Q.83 The average kinetic energy of the gas molecules is
 (a) Inversely proportional to its absolute temperature
 (b) Directly proportional to its absolute temperature
 (c) Equal to the square of its absolute temperature
 (d) Directly proportional to the square root of its absolute temperature
- Q.84 If the pressure and absolute temperature of 3 dm^3 of a gas are doubled, its volume would be
 (a) 2 dm^3 (b) 3 dm^3
 (c) 6 dm^3 (d) 12 dm^3
- Q.85 When small piece of sodium floated on water reacts vigorously to produce metal hydroxide and _____ gas
 (a) Oxygen (b) Nitrogen
 (c) Hydrogen (d) Both a and c



- Q.86 The compressibility factor, Z , for an ideal gas is
(a) Zero (b) Less than one
(c) Greater than one (d) Equal to one
- Q.87 The real gases show nearly ideal behaviour at
(a) Low pressure and low temperature (b) High pressure and low temperature
(c) High pressure and high temperature (d) Low pressure and high temperature
- Q.88 Which of the following is an example of molecular solid
(a) ZnS (b) MgO
(c) Ice (d) Graphite
- Q.89 Solid "X" melts over a wide range of temperature. On this basis, it is classified as
(a) Amorphous solid (b) Crystalline solid
(c) Molecular solid (d) Covalent solid
- Q.90 Which of the following is wrong?
(a) Water has maximum density at 4°C
(b) Molecular solids are relatively soft and volatile
(c) In diamond carbon atom has tetrahedral geometry
(d) Diamond is a good conductor of electricity
- Q.91 Which solid does not contain true covalent bonds?
(a) Aluminium nitride (b) Cadmium iodide
(c) Nickel (d) Graphite
- Q.92 Which order of mass is correct in ascending order
(a) $e^- > p^+ > n$ (b) $e^- < n < p^+$
(c) $e^- < p^+ < n$ (d) $n > p^+ > e^-$
- Q.93 A sub shell having $n = 6$, and $l = 3$ is designated as
(a) 5s (b) 6p
(c) 6f (d) 6d
- Q.94 Which of the following alkaline earth metal reacts with boiling water to form metal oxide and hydrogen gas
(a) Ca (b) Be
(c) Mg (d) Ba
- Q.95 In which of the following changes there is a transfer of five electrons
(a) $\text{MnO}_4^{2-} \longrightarrow \text{MnO}_2$ (b) $\text{CrO}_4^{2-} \longrightarrow \text{Cr}^{+3}$
(c) $\text{MnO}_4^- \longrightarrow \text{Mn}^{+2}$ (d) $\text{Cr}_2\text{O}_7^{2-} \longrightarrow 2\text{Cr}^{+3}$
- Q.96 The mass of 1×10^{-1} moles of MgSO_4 is
(a) 0.12g (b) 12.0g
(c) 1.2g (d) 120g
- Q.97 When gas is heated from 300 K to 400 K at a constant pressure of 1 atm, its volume
(a) Increases from V to 0.75V (b) Increases from V to 1.33V
(c) Increases from V to 1.5V (d) Increases from V to 2V
- Q.98 Conductivity of IIA group elements generally
(a) Increases (b) Remains same
(c) Decreases (d) No co-relation
- Q.99 The Boyle's law can be derived from ideal gas equation by holding:
(a) 'n' and 'P' constant (b) 'T' and 'V' constant
(c) 'n' and 'T' constant (d) 'P' and 'T' constant
- Q.100 Energy required to remove an electron from the outermost shell of its isolated gaseous atom in its ground state is:
(a) Electron affinity (b) Crystal energy
(c) Lattice energy (d) Ionization energy
- Q.101 The maximum number of electrons with clock wise spin that can be accommodated in "N" shell is
(a) 16 (b) 18
(c) 32 (d) 8
- Q.102 Which among the following have triple bonds
(I) NO_2 (II) HCN (III) CO (IV) CH_3CN
(a) Only (I), (II) and (III) (b) Only (I), (II) and (IV)
(c) Only (II), (III) and (IV) (d) Only (I), (III) and (IV)
- Q.103 Which one of the following is not a coplanar molecule
(a) HCN (b) C_2H_4
(c) SnCl_2 (d) C_2H_4



- Q.104 Heat absorbed by the system at constant pressure is equal to
(a) ΔE (b) ΔT
(c) ΔH (d) ΔP
- Q.105 Which of the following product is formed when ethyl chloride is treated with Na in presence of ether
(a) n-Butane (b) Ethane
(c) Propane (d) Methane
- Q.106 The alkyl halide which is most reactive
(a) CH_3F (b) CH_3Br
(c) CH_3Cl (d) CH_3I
- Q.107 The alcohol which readily gives nucleophilic substitution reaction
(a) Primary alcohols (b) Tertiary alcohols
(c) Secondary alcohols (d) All give equally
- Q.108 Alcohols have higher boiling points than their corresponding alkanes due to:
(a) Higher polarity (b) Ability to form hydrogen bonding
(c) Higher molar mass (d) All of these
- Q.109 The condition which will not help the fermentation of starch
(a) Proper aeration (b) Presence of preservative
(c) Dilution of solution (d) Optimum temperature 25°C to 35°C
- Q.110 Dehydration of alcohol produces alkenes. The alcohol which does not undergo this reaction is
(a) Methanol (b) 1-Propanol
(c) Ethanol (d) 2-Butanol
- Q.111 Which of the following will not show iodoform reaction?
(a) $\text{C}_6\text{H}_5\text{CHO}$ (b) CH_3COCH_3
(c) CH_3CHO (d) $\text{C}_6\text{H}_5\text{COCH}_3$
- Q.112 $\text{Na}_2\text{Cr}_2\text{O}_7$, what the product will be, when secondary alcohols are oxidized in same condition
(a) Alkenes (b) Alkynes
(c) Alkyl Halides (d) Ketones
- Q.113 Which of the following is more reactive
(a) Acetaldehyde (b) Acetone
(c) 2-Pentanone (d) Propionaldehyde
- Q.114 Which is not base catalyzed reaction of carbonyl compound like aldehyde
(a) Addition of NaHSO_3 (b) Haloform reaction
(c) Addition with HCN (d) Hydrazone formation
- Q.115 The compound obtained by reaction of phenol with an acid halide in the presence of a base is
(a) Salt (b) Ketone
(c) Sulphonic acid (d) Ester
- Q.116 $\text{CH}_3\text{CN} \xrightarrow{\text{H}_2\text{OH}^+} \text{B} \xrightarrow{\text{H}_2\text{OH}^+} \text{CH}_3\text{COOH} + \text{NH}_4^+$
'B' is
(a) $\text{CH}_3\text{CH}_2\text{OH} + \text{NH}_3$ (b) CH_3CONH_2
(c) $\text{CH}_3\text{CH}_2\text{NH}_2$ (d) $\text{CH}_3\text{CH}_2\text{NO}_2$
- Q.117 Acidic hydrolysis of propane nitrile produces
(a) Ethanoic acid (b) Acetamide
(c) Propanoic acid (d) Ethyl amine
- Q.118 Cannizzaro's reaction is not given by
(a) Formaldehyde (b) Benzaldehyde
(c) Acetaldehyde (d) Trimethylacetaldehyde
- Q.119 Number of covalent bonds in 3.6 g water
(a) 6.2×10^{24} (b) 2.4×10^{23}
(c) 1.2×10^{24} (d) 1.8×10^{24}
- Q.120 The compound which has same empirical formula as acetic acid
(a) Oxalic acid (b) Formic acid
(c) Sucrose (d) Glucose



MCQ'S RESPONSE FORM

ID	A	B	C	D	ID	A	B	C	D	ID	A	B	C	D	ID	A	B	C	D
1					52					104					156				
2					53					105					157				
3					54					106					158				
4					55					107					159				
5					56					108					160				
6					57					109					161				
7					58					110					162				
8					59					111					163				
9					60					112					164				
10					61					113					165				
11					62					114					166				
12					63					115					167				
13					64					116					168				
14					65					117					169				
15					66					118					170				
16					67					119					171				
17					68					120					172				
18					69					121					173				
19					70					122					174				
20					71					123					175				
21					72					124					176				
22					73					125					177				
23					74					126					178				
24					75					127					179				
25					76					128					180				
26					77					129					181				
27					78					130					182				
28					79					131					183				
29					80					132					184				
30					81					133					185				
31					82					134					186				
32					83					135					187				
33					84					136					188				
34					85					137					189				
35					86					138					190				
36					87					139					191				
37					88					140					192				
38					89					141					193				
39					90					142					194				
40					91					143					195				
41					92					144					196				
42					93					145					197				
43					94					146					198				
44					95					147					199				
45					96					148					200				
46					97					149									
47					98					150									
48					99					151									
49					100					152									
50					101					153									
51					102					154									
					103					155									

Roll No.									
0	0	0	0	0	0	0	0	0	0
1	1	1	1	1	1	1	1	1	1
2	2	2	2	2	2	2	2	2	2
3	3	3	3	3	3	3	3	3	3
4	4	4	4	4	4	4	4	4	4
5	5	5	5	5	5	5	5	5	5
6	6	6	6	6	6	6	6	6	6
7	7	7	7	7	7	7	7	7	7
8	8	8	8	8	8	8	8	8	8
9	9	9	9	9	9	9	9	9	9

NAME: _____

FATHER'S NAME: _____

ROLL NO. (IN WORDS): _____

CANDIDATE'S SIGNATURE: _____

DEPUTY SUPDT. SIGN: _____

INSTRUCTIONS

1. USE BLUE BALL POINT PEN ONLY.
2. PLEASE FILL IN THE ROLL NO. CORRECTLY.
3. IT IS IMPORTANT THAT THE CIRCLE IS FILLED COMPLETELY AND CORRECTLY AS SHOWN IN THE EXAMPLE BELOW, OTHERWISE THE UNIVERSITY CAN NOT BE HELD RESPONSIBLE.

CORRECT EXAMPLE: ○ ● ○ ○ ✓

INCORRECT EXAMPLES: ○ ● ● ○ X
○ ○ ● ○ X
○ ○ ○ ● X

4. DO NOT ERASE A RESPONSE ONCE THE CIRCLE HAS BEEN FILLED IN.
5. INCOMPLETELY FILLED CIRCLES WILL NOT BE READ.
6. MULTIPLE RESPONSE TO ONE QUESTION IS NOT ALLOWED.
7. TEARING OFF THE RESPONSE FORM, FOLDING, STAPLING, CUTTING & PUTTING UNNECESSARY SIGNS AND IDENTIFICATION ON THE FORM WILL LEAD TO AUTOMATIC DISQUALIFICATION OF THE CANDIDATE.

THE UNIVERSITY SHALL NOT BE HELD RESPONSIBLE IF THE ABOVE INSTRUCTIONS ARE NOT FOLLOWED.